**Chapter
Test C****Elements and Their Properties****I. Testing Concepts**

Directions: *In the blank, write the term that best completes each statement or answers the question.*

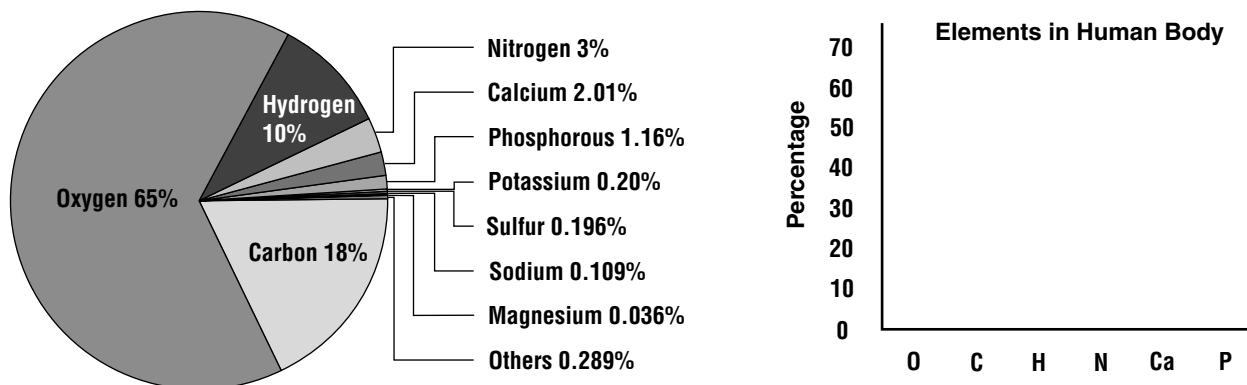
- _____ is the only metal that is not solid at room temperature.
- When metals combine with nonmetals the atoms of the metals tend to lose electrons, forming _____ bonds.
- In _____ bonding, positively charged metallic ions are surrounded by a cloud of electrons.
- A(n) _____ element is one in which the nucleus breaks down and gives off particles and energy.
- Elements in Groups 3 through 12 on the periodic table are known as _____ elements.
- The two rows of elements that seem to be disconnected from the rest of the periodic table are the _____.
- Are metals poor or good conductors of electricity? _____
- The elements in Group 1 of the periodic table are the _____ metals.
- Iron, cobalt, and nickel are known as _____.
- Copper, silver, and gold are known as _____.
- Metals in Earth's _____ must be mined and separated from their ores.
- Most of your body's mass is made of _____.
- A(n) _____ molecule consists of two atoms of the same element in a covalent bond.
- If a halogen gains an electron from a metal, a(n) _____ is formed.
- Are the noble gases stable or unstable? _____
- _____ can form ionic and covalent bonds with other elements and can have metallic and nonmetallic properties.
- Different forms of the same element are called _____.
- Elements having more than 92 protons are called _____ elements.

Chapter Test C (continued)

II. Understanding Concepts

Skill: Making and Using Graphs

Directions: Use the pie graph below to answer questions 1 through 3.



- Use the percentages shown in the pie graph to complete the bar chart above. Show the percentage for the six elements that make up most of the human body.
- Which three elements make up more than 90 percent of the human body?

- If your mass is 50 kg, how much of your body is nitrogen?

Skill: Interpreting Scientific Illustrations

Directions: In the space below, draw a hydrogen atom. Label the nucleus with one proton and zero neutrons. Draw an electron orbiting the nucleus in the first energy level. Then draw a lithium atom. Label the nucleus with three protons and three neutrons. Then draw two electrons orbiting the nucleus in the first energy level and one electron orbiting the nucleus in the second energy level.

- How are the hydrogen atom and the lithium atom similar?

- What is the atomic number of hydrogen? Of lithium?

- Why is hydrogen placed in the same group as the alkali metals?
